Potential significance of photoexcited NO$_2$ on global air quality with the NMMB/BSC chemical transport model


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Atmospheric chemists have recently focused on the relevance of the NO$_2^+$ + H$_2$O → OH + HONO reaction to local air quality. This chemistry has been considered not relevant for the troposphere from known reaction rates until nowadays. New experiments suggested a rate constant of $1.7 \times 10^{-13}$ cm$^3$ molecule$^{-1}$ s$^{-1}$, which is an order of magnitude faster than the previously estimated upper limit of $1.2 \times 10^{-14}$ cm$^3$ molecule$^{-1}$ s$^{-1}$, determined by Crowley and Carl (1997). Using the new global model, NMMB/BSC Chemical Transport Model (NMMB/BSC-CTM), simulations are presented that assess the potential significance of this chemistry on global air quality. Results show that if the NO$_2^+$ chemistry is considered following the upper limit kinetics recommended by Crowley and Carl (1997), it produces an enhancement of ozone surface concentrations of 4–6 ppbv in urban areas and 6–15 ppbv in urban locations, reaching a maximum enhancement of 30 ppbv in eastern Asia. Moreover, NO$_2$ enhancements are minor (<0.01 ppbv) in background regions and reach maximum daytime values of 1–6 ppbv. Similarly, HONO exhibits negligible increases, 8–9 pptv in urban settings. Enhancements in the concentration of OH are around 14–17 × 10$^3$ molec cm$^{-3}$. Decreases in the concentration of O$_3$ and its precursors are also identified but to a lesser degree. In order to quantify the role of the two kinetic rates measured, model simulations are compared after incorporating both reaction rate constants. Maximum O$_3$ difference enhancements from 5 to 10 ppbv are modeled over locations where high NO$_x$ emissions are present; however, differences are small in most parts of the globe.


1. Introduction

The chemistry of electronically excited NO$_2$ and H$_2$O has gained interest during the last years for its implication in the enhancement of hydroxyl radical (OH) production. It is well-known that OH radicals are the most important oxidant in the atmosphere during the day. Photolysis of O$_3$ in the presence of water vapor is the major source of OH radicals:

$$O_3 + h\nu(\lambda \leq 320\text{nm}) \rightarrow O_2 + O(^{1}\text{D}) \quad (1)$$

$$O(^1\text{D}) + H_2O \rightarrow 2OH \quad (2)$$

$$O(^1\text{D}) + M \rightarrow O(^3\text{P}) + M. \quad (3)$$

In addition, other well-recognized sources of OH in the troposphere are

$$\text{HONO} + h\nu(\lambda < 390\text{nm}) \rightarrow \text{OH} + \text{NO} \quad (4)$$

$$\text{H}_2\text{O}_2 + h\nu \rightarrow 2\text{OH} \quad (5)$$

$$\text{HO}_2 + \text{NO} \rightarrow \text{NO}_2 + \text{OH}. \quad (6)$$

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Fast gas-phase photochemical cycles involving NO$_2$ influence tropospheric O$_3$ concentrations. NO$_2$ absorbs sunlight strongly and is in rapid photochemical equilibrium with NO and O$_3$ [Leighton, 1961]:

$$\text{NO}_2 + h\nu (\lambda \leq 420\text{nm}) \rightarrow \text{NO} + \text{O} \quad (7)$$

$$\text{O} + \text{O}_2 + \text{M} \rightarrow \text{O}_3 + \text{M} \quad (8)$$

$$\text{NO} + \text{O}_3 \rightarrow \text{NO}_2 + \text{O}_2. \quad (9)$$

The NO$_2$ absorption spectrum displays two broad continua centered at ca. 400 and 210 nm [Crowley and Carl, 1997]. Due to the filtering effect of O$_3$ in the stratosphere, only wavelengths longer than 300 nm are relevant for photochemistry in the troposphere. Under high tropospheric solar zenith angles, the UV flux is depleted, and other OH production mechanisms become relevant. In this sense, when NO$_2$ absorbs sunlight between 400 and 650 nm, NO$_2$ photoexcites to NO$^*$. Then, the reaction of NO$^*$ with H$_2$O can be a major source of tropospheric OH radicals [Li et al., 2008]:

$$\text{NO}_2 + h\nu (\lambda > 420\text{nm}) \rightarrow \text{NO}_2^* \quad (10)$$

$$\text{NO}_2^* + \text{M} \rightarrow \text{NO}_2 + \text{M} \quad (11)$$

$$\text{NO}_2^* + \text{H}_2\text{O} \rightarrow \text{OH} + \text{HONO}. \quad (12)$$

Li et al. [2008] reported that the bimolecular reaction of electronically excited NO$^*$ with H$_2$O can lead to substantial OH production. Previous results from Crowley and Carl [1997] suggested that the photoexcitation chemistry of NO$_2$ producing OH radicals is not relevant in the atmosphere, suggesting an upper limit rate constant for the photoexcited NO$_2^*$ with H$_2$O reaction of $1.2 \times 10^{-14}$ cm$^3$ molecule$^{-1}$ s$^{-1}$, an order of magnitude lower than that determined by Li et al. [2008], $1.7 \times 10^{-13}$ cm$^3$ molecule$^{-1}$ s$^{-1}$.

Crowley and Carl [1997] and Li et al. [2008] add another level of complexity to the current understanding of production of ozone and its precursors. Reaction rate constants quantified by these studies, therefore, lead to enhanced production of O$_3$ due to the net production of OH. This is due to the comparable reaction efficiency of OH production via the NO$_2^*$ + H$_2$O reaction channel (e.g., 1 in 10,000 of the NO$_2^*$ molecules produced react with H$_2$O to produce OH and HONO) compared as described in reactions (1–3) to the classical OH source (i.e., by way of ozone photolysis) and the fact that the amount of OH produced from NO$_2^*$ scales linearly with the amount of NO$_2$ in the atmosphere [Wennberg and Dabdub, 2008]. In other words, despite the minor branching of the NO$_2 + h\nu$ reaction to produce NO$^*$, which in turn reacts with H$_2$O to produces OH and HONO, typical concentrations of NO$_2$ (30 to 100 ppbv = $7 \times 10^{11}$ to $2 \times 10^{12}$ molecule cm$^{-3}$) are much larger than that of OH ($8 \times 10^{-3}$ to 0.5 pptv = $2 \times 10^5$ to $1 \times 10^7$ molecule cm$^{-3}$).

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Additional studies have investigated the impact of reaction (12) [Wentzell et al., 2010; Li et al., 2011; Sörgel et al., 2011]. Wentzell et al. [2010] modeled observed HONO values at the Harrow, Ontario, Canada supersite, via incorporating reaction (12) and could not account for the measured excess HONO produced since the calculated production rate is a factor of 2–6 larger. Sörgel et al. [2011] utilized the Crowley and Carl [1997] and Li et al. [2008] kinetics to model HONO measurements during the Diel Oxidant Mechanism In relation to Nitrogen Oxides (DOMINO) campaign in southwest Spain; they showed that the Li et al. [2008] kinetics contributed to ~8% toward HONO measurements, while using the Crowley and Carl [1997] kinetics exhibited negligible amounts of HONO.

Li et al. [2011] used the WRF-Chem model to model HONO, NO$_x$, and O$_3$ observations over Beijing, Tianjin, and Hebei Provinces of China, which they note (in addition, to including updated HONO emissions, heterogeneous reactions on aerosol surfaces), considerably improve HONO and O$_3$ simulations in these respective regions.

It has been shown that the emission regime plays an important role on the impacts of the chemistry [e.g., Wennberg and Dabdub, 2008; Sarwar et al., 2009; Ensberg et al., 2010]. From all previous results, an important question arises: What is the possible impact of photoexcited NO$_2$ chemistry on a global scale? In the present work, global simulations, produced from an online chemical transport model and constrained by reaction rate constants proposed by Crowley and Carl [1997] (as an upper limit) and Li et al. [2008] are analyzed for the first time to elucidate
quantitatively the impact of the NO\textsuperscript{\textit{P}} chemistry. The impact of the photoexcited chemistry is first discussed within the context of comparing simulations with and without Crowley and Carl [1997] NO\textsuperscript{\textit{P}} kinetics within a default tropospheric chemical mechanism, Carbon-Bond 05 (CB05) [Tarwood et al., 2005]. Thereafter, the impact of considering Li et al. [2008] NO\textsuperscript{\textit{P}} kinetics is also evaluated in comparison to both simulations, with and without Crowley and Carl [1997] NO\textsuperscript{\textit{P}} kinetics.

[11] The following section discusses the uncertainties among proposed reaction rate constants for reaction (12). Section 3 describes the model and the experimental setup. The results and evaluation of the model simulations are discussed in section 4, and a final summary and conclusion section is presented in section 5.

2. Uncertainties on Reaction Rates of Photoexcited NO\textsuperscript{\textit{P}}

[12] Li et al. [2008] reported a reaction rate constant of \( k_{12} = 1.7 \times 10^{-13} \) cm\(^2\) molecule\(^{-1}\) s\(^{-1}\), which is an order magnitude greater than the upper limit proposed by Crowley and Carl [1997] (\( k_{12} = 1.2 \times 10^{-14} \) cm\(^2\) molecule\(^{-1}\) s\(^{-1}\)). Thereafter, Carr et al. [2009] argued that the NO\textsuperscript{\textit{P}} reaction with H\textsubscript{2}O likely does not occur under normal atmospheric conditions although they produce a reaction rate a factor of 17 lower than that reported by Li et al. [2008], which similar to the results of Crowley and Carl [1997] at the 2 \( \sigma \) level. Similar to Carr et al. [2009], Li et al. [2009] then responded by stating that the reason for the differences in the results of the two studies are not obvious especially when considering that the reaction under consideration is a difficult one to study due to rapid quenching of NO\textsuperscript{\textit{P}} by water and low product yields [Li et al., 2008, Carr et al. [2009] and Li et al. [2009] both point out that the difference in experimental approach may contribute to this incongruence, specifically, the use of unfocused laser light [Carr et al., 2009] compared to the use of a lens to focus the laser light [Li et al., 2008]. Li et al. [2008] used OH from the vibrational overtone-induced unimolecular dissociation of CH\textsubscript{3}OOH (S\textsubscript{5\textit{v(\textit{o})}}) to calibrate OH yields, whereas Carr et al. [2009] used the photodissociation of acetone, followed by the reaction of the resulting acetyl radical with O\textsubscript{2} for their calibrations. Carr et al. [2009] suggest that the differences in the results may be primarily due to multiphoton effects, contributing to the OH signal in their study, which would be sourced from the greater laser power density created by the presence of the lens. Carr et al. [2009] cites the presence of an intercept appearing the linear power dependence plot in Li et al. [2008] as indicative of multiphoton effects. Carr et al. [2009] did not suggest a mechanism for which OH could be produced by the multiphoton excitation process and was not able to provide data on the effect of increasing power density in their measurements of the OH yield as the maximum laser fluence used their experiments was below the threshold required by Li et al. [2008] to generate an OH signal. Yet Li et al. [2009] found no compelling evidence that multiphoton effects contributed to formation of OH radicals after investigating several multiphoton mechanism: a) multiphoton dissociation of NO\textsubscript{2} to produce O(\textit{D}), which can react with H\textsubscript{2} to produce OH radicals [Sander et al., 2006]; b) the production of translationally hot ground state O(\textit{P}) atoms from the multiphoton dissociation of NO\textsubscript{2}, which in principle can react with water to produce OH radicals; c) populating long-lived bound excited quartet electronic excited states of NO\textsubscript{2} through multiphoton excitation, which then reacts with H\textsubscript{2}O to produce OH, and d) the primary source of measured OH signal were due to the collision of electronically excited NO\textsubscript{2} with vibrationally excited H\textsubscript{2}O. The first scenario was ruled out as the introduction of H\textsubscript{2}O in to the experimental cell in place of H\textsubscript{2}O did not produce any measurable amount of OH. The second scenario was also ruled out as the O(\textit{P}) reaction is endothermic by \( \sim 17 \) kcal mole\(^{-1}\) [Sander et al., 2006], and the translational energy generated from the photodissociation of NO\textsubscript{2} is not sufficiently energetic to initiate the O(\textit{P}) + H\textsubscript{2}O reaction (e.g., 0.3 eV (\( \sim 6.9 \) kcal mole\(^{-1}\) at 308 nm)). Although NO\textsubscript{2} has several excited quartet states in the vicinity of 3.61 to 4.37 eV, accessing these bound excited electronic NO\textsubscript{2} from the ground state requires the involvement of spin-forbidden transitions. Moreover, although these bound excited states could potentially be accessed through sequential two-photon excitation by the intermediate A\textsuperscript{\textit{B}}\textsubscript{2} electronic state of NO\textsubscript{2}, initially populated through the excitation in the visible range, spectral features of bound-to-bound transitions, such as these should be reflected in their characteristic structured excitation spectrum. Li et al. [2008] produce an action spectrum from the NO\textsubscript{2} + H\textsubscript{2}O reaction that match the expected X\textsuperscript{X}\textsubscript{A} \( \rightarrow \) A\textsuperscript{A}\textsubscript{B} excitation of NO\textsubscript{2} in the visible range. In addition to being spin-forbidden in regards to their photochemical formation, the bimolecular reaction of bound excited state NO\textsubscript{2} with H\textsubscript{2}O is also expected to be spin-forbidden. The last scenario was ruled out as Li et al. [2008] did not produce a squared-dependence for the pseudo first-order rates. They do note, however, that this pathway, may contribute to small background levels of OH.

[13] Recently, Amedro et al. [2011] contribute to this controversial reaction by way of using the coupling of a Fluorescence Assay be Gas Expansion (FAGE) instrument to a laser photolysis cell. They claim that no OH radicals were observed using an unfocused excitation laser beam at 565 nm, but OH was measured by focusing the laser beam. Based on the dependence of the OH-fluorescence signal with the laser energy, they deduce that the OH radicals are sourced from a complex mechanism that includes multiphoton absorption of NO\textsubscript{2}. Amedro et al. [2011] quantify an upper-limit OH yield of \( 8 \times 10^{-6} \) for reaction (12), which is in accordance with Crowley and Carl [1997] and Carr et al. [2009], who estimate OH-yield upper-limits of \( 7 \times 10^{-5} \) and \( 6 \times 10^{-5} \), respectively. Within this context, Crowley and Carl [1997] also quantify a reaction rate for reaction (12) of \( 1.2 \times 10^{-15} \) cm\(^3\) molecule\(^{-1}\) s\(^{-1}\), which is, again, 1-order of magnitude less than that measured by Li et al. [2008], i.e., \( 1.7 \times 10^{-15} \) cm\(^3\) molecule\(^{-1}\) s\(^{-1}\).

[14] Assuming all the uncertainties around reaction (12) rate constant, in the present work, the potential significance of this chemistry is assessed at global scale applying the suggested rates of both Li et al. [2008] and Crowley and Carl [1997].

3. Methods

3.1. Numerical Model

[15] The model used for this study is the NMMB/BSC Chemical Transport Model (NMMB/BSC-CTM) [Pérez
Table 1. Model Configuration

<table>
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<tr>
<th>Meteorology</th>
<th>Configuration</th>
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<tr>
<td>Dynamics</td>
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<td>Binkowsky scheme</td>
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<td>MEGAN</td>
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<td>Resolution and Initial Conditions</td>
<td>Value</td>
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<tr>
<td>Vertical layers</td>
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<tr>
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<tr>
<td>Initial condition</td>
<td>LMDz-INCA</td>
</tr>
<tr>
<td>Spin-up</td>
<td>6 month</td>
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</tbody>
</table>

et al., 2011; Jorba et al., 2011; Pérez et al., 2009; Jorba et al., 2009). It is an online chemical weather prediction system for mesoscale to global-scale applications, developed at the Barcelona Supercomputing Center. The atmospheric driver is the NCEP/NMMB multiscale numerical weather prediction model, developed at the National Centers for Environmental Prediction (NCEP). A description of the modeling system is presented in this section. Table 1 summarizes the main characteristics of the NMMB/BSC-CTM model.

3.1.1. The Atmospheric Driver: NMMB

[16] The Non-hydrostatic Multiscale Model on the B grid (NMMB) [Janjic, 2005; Janjic and Black, 2007; Janjic et al., 2011] is a new unified atmospheric model for a broad range of spatial and temporal scales. Its unified non-hydrostatic dynamical core allows regional and global simulations. The regional NMMB is the operational regional North American Mesoscale (NAM) model at NCEP. The numerical schemes used in the model were designed following the principles set up in Janjic [1977, 1979, 1984, 2003]. Isotropic horizontal finite volume differencing is employed so a variety of basic and derived dynamical and quadratic quantities are conserved. The hybrid pressure-sigma coordinate [Simmons and Burridge, 1981] is used in the vertical and the Arakawa B-grid horizontal staggering is applied in the horizontal. The unified version of the model is developed for the scales ranging from large eddy simulations (LES) to global simulations [Janjic, 2005]. The non-hydrostatic component of the model dynamics is introduced through an add-on module that can be turned on or off, depending on resolution. The physical package used is composed by: (1) the Mellor-Yamada-Janjic (MYJ) level 2.5 turbulence closure for the treatment of turbulence in the planetary boundary layer (PBL) and in the free atmosphere [Janjic, 2001], (2) the surface layer scheme based on the Monin-Obukhov similarity theory [Monin and Obukhov, 1954] with introduced viscous sublayer over land and water [Zilitinkevich, 1965; Janjic, 1994], (3) the NCEP NOAH land surface model [Ek et al., 2003] or the LISS model [Vukovic et al., 2010] for the computation of the heat and moisture surface fluxes, (4) the GFDL longwave and shortwave radiation [Fels and Schwarzkopf, 1975; Lacis and Hansen, 1974], (5) the Ferrier gridscale clouds and microphysics [Ferrier et al., 2002], and (6) the Betts-Miller-Janjic convective adjustment scheme [Betts, 1986; Betts and Miller, 1986; Janjic, 1994, 2000]. Vertical diffusion is handled by the surface layer scheme and by the PBL scheme. Lateral diffusion is formulated following the Smagorinsky non-linear approach [Janjic, 1990].

3.1.2. Online Gas-Phase Chemistry

[17] The gas-phase tropospheric chemistry is inlined within the NMMB code. Meteorological information is available at each time step to solve the chemical processes properly. In order to maintain full consistency with the meteorological solver, chemical species are advected and mixed at the corresponding time step of the meteorological tracers using the same numerical schemes implemented in the NMMB. The advection scheme is Eulerian, positive definite and monotone, maintaining a consistent mass-conservation of the chemical species within the domain of study [Janjic et al., 2009; Tang et al., 2009].

[18] For the present study, the Carbon-Bond 2005 (CB05) is used [Yarwood et al., 2005]. It is an updated version of the Carbon-Bond IV (CB4) lumped-structure-type mechanism of Gery et al. [1989]. The CB4 mechanism was formulated focusing on limited domain extent, urban environments and for planetary boundary layer chemistry. In this sense, the CB05 extends its applicability from urban to remote tropospheric conditions and is suitable for global applications.

[19] The CB05 core mechanism has 51 chemical species and solves 156 reactions. It considers 20 photolytic reactions and extends inorganic reactions from urban to remote tropospheric conditions. The rate constants are updated based on evaluations of Atkinson et al. [2004] and Sander et al. [2006]. The organic compounds not explicitly treated are apportioned to the carbon-bond species based on the molecular structure and following Yarwood et al. [2005] assignments from VOC species to CB05 model species.

[20] No explicit stratospheric chemistry is implemented in the current version of the model. Stratospheric gas-phase concentrations are initialized from climatological profiles or from a global chemical model and are constant throughout a given simulation.

[21] One of the most important processes determining tropospheric composition is the photo-dissociation of trace gases. 20 photolysis reactions are considered in the standard configuration of the model. To compute the photolysis rates, the Fast-J [Wild et al., 2000] online photolysis scheme is implemented. The quantum yields and cross section for the CB05 photolysis reactions have been revised and updated following the recommendations of Atkinson et al. [2004] and Sander et al. [2006].

[22] An online implementation is used to characterize the dry-deposition of atmospheric species following the Wesely [1989] scheme. The cloud chemistry scheme of Byun and Ching [1999] and Foley et al. [2010] has been implemented within NMMB/BSC-CTM in order to resolve the cloud processes affecting the concentration of air pollutants. The processes included are grid-scale scavenging and wet-deposition, subgrid-scale vertical mixing, scavenging and wet-deposition for precipitating and non-precipitating clouds. Only in-cloud scavenging is considered in the current implementation.
Concerning natural on-line emissions, only biogenic emissions are considered as natural source. Biogenic emissions are computed on-line from the Model of Emissions of Gases and Aerosols from Nature (MEGAN) [Guenther et al., 2006]. MEGAN estimates the emission of more than 130 non-methane volatile organic compounds (NMVOCs). All the MEGAN NMVOCs are speciated, following the CB05 chemical mechanism; thus, emissions for isoprene, lumped terpenes, methanol, nitrogen monoxide, acetaldehyde, ethanol, formaldehyde, higher aldehydes, toluene, carbon monoxide, ethane, ethene and paraffin carbon bond, and olefin carbon bond are considered within the chemical processes of the NMMB/BSC-CTM model. Biogenic emissions are computed every hour, accounting for the meteorological evolution changes of solar radiation, surface temperature, moisture and precipitation. Figure 1a shows the daily mean emission of isoprene for July 2004 over the global domain.

3.2. Global Anthropogenic Emissions

Global anthropogenic emissions applied in the present study are based on the Emmons et al. [2010] emission database. The majority of the anthropogenic emissions originated from the POET (Precursors of Ozone and their Effects in the Troposphere) database for 2000 (C. Granier et al., POET, a
Table 2. Model Runs

<table>
<thead>
<tr>
<th>Case Name</th>
<th>Abbreviation</th>
<th>Description</th>
</tr>
</thead>
<tbody>
<tr>
<td>Base case</td>
<td>BC</td>
<td>NMMB/BSC-CTM model</td>
</tr>
<tr>
<td>Crowley and Carl case</td>
<td>CC</td>
<td>NMMB/BSC-CTM model with default CB05 chemistry and the photoexcitation chemistry of NO₂ with the kinetics of Crowley and Carl [1997]</td>
</tr>
<tr>
<td>Li et al. case</td>
<td>LC</td>
<td>NMMB/BSC-CTM model with default CB05 chemistry and the photoexcitation chemistry of NO₂ with the kinetics of Li et al. [2008]</td>
</tr>
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</table>

3.3. Implementation of the NO₂⁺ Chemistry

[25] The photoexcitation of NO₂⁺ (reactions (10)-(12)) has been implemented within the CB05 chemical mechanism. The kinetic constant used for reaction (11) is \( k_{11} = 2.8 \times 10^{-11} \text{ cm}^3 \text{ molecule}^{-1} \text{s}^{-1} \) [Okabe, 1978; Crowley and Carl, 1997]. For reaction (12), two options are implemented: (a) the constant of Li et al. [2008] \( k_{12} = 1.7 \times 10^{-13} \text{ cm}^3 \text{ molecule}^{-1} \text{s}^{-1} \), and (b) the Crowley and Carl [1997] kinetic constant upper limit \( k_{12} = 1.2 \times 10^{-14} \text{ cm}^3 \text{ molecule}^{-1} \text{s}^{-1} \). Thus, three different cases are prepared: a default CB05 without the photoexcitation of NO₂⁺, a CB05 with fast kinetics from Li et al. [2008] and a last mechanism with a lower reaction rate constant from Crowley and Carl [1997].

[26] In order to consider the photoexcitation of NO₂⁺, the photolysis rate of reaction (10) is computed within the Fast-J photolysis scheme. The quantum yields and cross sections of NO₂⁺ have been implemented within the NMMB/BSC-CTM model. Results at the first surface model layer for a specific time step show a maximum value for the photolysis rate of \( J_{\text{NO}_2} = 0.6 \text{ min}^{-1} \) and a maximum value for the photoexcitation rate of \( J_{\text{NO}_2^+} = 1.6 \text{ min}^{-1} \).

[28] Additionally, dry-deposition of NO₂⁺ has also been considered. The approach used is that of Wesely [1989] for NO₂. However, tests show that the reaction is fast enough to consider the NO₂⁺ dry-deposition negligible (not shown).

3.4. Description of the Model Runs

[29] In order to analyze the impact of the NO₂⁺ chemistry on tropospheric ozone concentration levels at global scale, three model configurations are defined (see Table 2): a Base Case run (BC) with the default CB05 chemical mechanism, a Crowley and Carl [1997] Case run (CC) with \( k_{12} = 1.2 \times 10^{-14} \text{ cm}^3 \text{ molecule}^{-1} \text{s}^{-1} \), and a Li et al. [2008] Case run (LC) with \( k_{12} = 1.7 \times 10^{-13} \text{ cm}^3 \text{ molecule}^{-1} \text{s}^{-1} \). The NMMB/BSC-CTM is configured at a global scale with the dynamics, physics and chemistry described in Table 1.

[30] Thus, three model runs have been performed (BC, CC and LC) for July and August 2004. The horizontal resolution of the global model is \( 1^\circ \times 1.4^\circ \) and uses 64 vertical sigma-hybrid layers from surface to 0 hPa. The meteorological initial conditions are obtained from the final analysis of NCEP global model for 2004 at 1° horizontal resolution. A spin-up of 6 months was run to prepare the chemical conditions for the July and August 2004 runs. The chemical initial conditions of the spin-up were obtained from the LMDz-INCA2 model [Szopa et al., 2009]. The meteorology is run 24 h and reinitialized each day with the corresponding analysis, while the chemical initial conditions of the next day are those from the previous day model run at 24 h.

3.5. Network of Surface Observations

[31] Background surface observations of O₃, NO₂ and CO are used to evaluate the results of the model NMMB/BSC-CTM. Stations from the World Data Centre for Greenhouse Gases (WDCGG; http://gaw.kishou.go.jp/wdcgg/) have been selected for the evaluation. The WDCGG stations measure O₃, NO₂ and CO in background areas. Moreover, stations from the Co-operative Programme for Monitoring and Evaluation of the Long-range Transmission of Air Pollutants in Europe (EMEP; http://www.emep.int/) have been selected to complement the observations of WDCGG over Europe, and the Clean Air Status and Trends Network (CASTNET; http://java.epa.gov/castnet/) complements the observations over United States of America. Observations of surface O₃ and NO₂ are selected from the EMEP network, while only O₃ observations are obtained from the CASTNET network. The total number of stations used in the evaluation is indicated in Table 5.

[32] Classical statistics are selected to evaluate the model and quantify the impact of the NO₂⁺ on global air quality. According to the recommendations of the US-EPA a cut-off value of 80 \( \mu \text{g} \text{m}^{-3} \) is applied to O₃ statistics. The monthly mean model-to-data statistics for the mean bias (MB), root mean square error (RMSE), mean normalized bias error (MNBE), mean normalized gross error (MNGE), mean fractional bias (MFB) and mean fractional error (MFE) are
selected to provide quantitative scores about the performance of the model and the impact of the new photoexcited pathway.

4. Results and Discussion

[33] The impact of the NO$_2$ photoexcitation chemistry is presented in the following subsections. Specifically, the BC and CC model runs are compared and discussed during July August 2004. These discussions are focused on the northern hemisphere, where major anthropogenic emissions are sourced (see Figure 1b); in addition, this dialog contextualizes recent investigations [Sarwar et al., 2009; Ensb erg et al., 2010], that also focused on the northern hemisphere. In addition, LC and CC simulations are also discussed. Furthermore, given the uncertainties inherent to the reaction rate obtained by Li et al. [2008] and Carr et al. [2009], the impact of the NO$_2^+$ chemistry is assessed by comparing the BC and CC runs.

4.1. Impact of NO$_2$ Photoexcitation on Surface Ozone

[34] Figure 2 shows the hourly, 8-h average, and daily mean maximum ozone enhancement at surface level when comparing the CC results with the BC results. Results show how the CC chemistry contributes to a maximum increase in the hourly ozone surface concentration around 10–15 ppbv in local urban areas, reaching a maximum of 31 ppbv in Taiwan, and approximately 3–7 ppbv in rural zones. Figure 2b shows the 8-h average maximum enhancements. When compared with the maximum hourly data, it clearly appears a pattern centered on local polluted regions, where the hourly enhancement is about 5 ppbv higher. The new chemistry significantly increases ozone concentrations during its peak hours of formation, which is primarily due to an increase in the production rate of OH radicals. Additionally, the 8-h average concentrations increase 6–8 ppbv in regional continental areas and 8–12 ppbv over water bodies. Major enhancements are observed again in large polluted regions (e.g., southern California area, Florida western coast, northwestern European coast, northern Taiwan, and east China). The daily mean maximum enhancements observed during July and August 2004 (Figure 2c) remain below 7 ppbv over most continental areas of the northern hemisphere. Only in Taiwan, the impact is more relevant and ozone daily mean concentrations may increase more than 15 ppbv.

[35] The areas where the impact of the NO$_2$ photoexcitation chemistry appears to be relevant are over ocean and sea in agreement with Sarwar [2009]. The limited ozone sink processes over water bodies, the shallow marine boundary layer, and the effectiveness of the new rate constant (i.e., by increasing OH production rate) promote higher ozone concentrations. Furthermore, those areas with larger enhancement are characterized by a large NO$_2$/VOC ratio. Figure 3 shows the maximum value of the ratio NO$_2$/VOC for July and August at ground level. Indeed, the ship routes are characterized by large NO$_2$/VOC ratios and strongly contribute to the enhancement of O$_3$ observed over the oceans. Additionally, enriched NO$_2$ plumes advected from continental areas toward the oceans provides efficient conditions where NO$_2^+$ chemistry plays an important contribution in the background ozone concentrations. Specifically, peroxo acetyl nitrate (PAN) is enhanced when the photoexcitation chemistry is activated. The increase is larger over oceans than over continental areas (not shown), suggesting that the long-range transport of PAN is also an important process that contributes to the enhancement of O$_3$ over the oceans.

[36] Additionally, atmospheric dynamics over the different ocean areas strongly contributes to the occurrence of ozone enhancements. For instance, the northern Atlantic ocean presents wide areas where ozone enhancements at surface are over 6 ppbv. This behavior is also observed in the northern Pacific ocean, but in this case, the extension where the enhancements reach values over 6 ppbv are not as large as those predicted over the Atlantic. Thus, the impact over continental areas due to the advection of ozone enriched air is more important in Europe (especially northern Europe) than in western USA. While the latter shows important contributions of local emissions, in Europe the background enhanced concentrations overwhelm the local enhancements due to the NO$_2^+$ chemistry. Of significant relevance are the enhancements over the Mexican gulf, the Mediterranean sea, the Persian gulf and the East China sea. Those are areas where atmospheric dynamics play an important role on the accumulation of aged air masses and are affected by large emission sources nearby.

[37] The temporal evolution of the spatially averaged daily mean ozone enhancement over the northern hemisphere, USA, northern Atlantic ocean, Europe, eastern Asia, and the northern Pacific ocean (see Figure 2c) are computed and plotted for the two months analyzed (Figure 4). There is a clear increasing trend in the daily mean enhancements from July to August. In the northern hemisphere, the ozone enhancements go from 2.5 ppbv in the beginning of July to 3 ppbv at the end of August. Such a trend is directly related to the summer climate conditions. The areas where the impact is more marked are the north Atlantic ocean and the north Pacific ocean, where the mean enhancements ranges from 2 to 2.5 ppbv during the two months. There are important increases in ozone enhancements (e.g., during the first week of August in the Pacific ocean) over the oceans. Those are attributed to episodes of strong advection of continental air masses enriched with anthropogenic pollutants. Over continental areas, the enhancement is around 2 ppbv over Europe and USA, and 1.5 ppbv in eastern Asia. Furthermore, Table 3 summarizes the ozone enhancements observed over the different areas. Maximum spatial-averaged enhancements range from 2.6 in the North Atlantic to 1.6 over Asia. It is important to note that also ozone decreases are produced when comparing CC and BC. In this sense, USA, Europe and Asia show ozone decreases of 3.8, 1.5 and 2.3 ppbv respectively. However, the mean behavior in all areas involves an enhancement of O$_3$.

[38] In order to complement the maximum O$_3$ enhancements analysis, the number of days where a threshold in the ozone enhancement is exceeded are computed. Table 4 contains the spatially average number of days per region (i.e., North Pacific, USA, North Atlantic, Europe, Asia (see Figure 2c)), and the maximum number of days within a region where an enhancement greater than or equal to 3, 5, 10 or 15 ppbv is modeled. In all areas analyzed, there is always a location that exceeds the 3 ppbv threshold. The North Pacific and USA areas are the ones with more exceedances (up to
Figure 2. (a) Hourly, (b) 8-h average, and (c) daily mean maximum ozone enhancement at ground level for July and August 2004 due to the inclusion of Crowley and Carl [1997] NO$_2$ photoexcitation chemistry. [Lower panel shows the location of the zonal averages for NAtl (red square), EU (dash-dotted black square), Asia (dashed black square), NPacific (dashed red square) and USA (black square) zones].
26 days the threshold of 10 ppbv is exceeded) than the rest of the north hemisphere. The North Atlantic, Europe and Asia zones present the lowest number of days with high exceedances (10 and 15 ppbv).

Compared with other studies focused on the NO$_2$ chemistry reported over USA, Sarwar et al. [2009] showed a maximum increase in O$_3$ concentrations of 9 ppbv in eastern USA and 6 ppbv in western USA for July 2001 and 2002, and Ensberg et al. [2010] reported maximum increases in the 8-h average O$_3$ concentrations of 8 ppbv for the South Coast Air Basin of California for summer conditions of 2005. Results are in agreement with those previous simulations. However, the impact of the new chemistry is more relevant downwind of important polluted areas, and there is an important effect on the long-range transport chemistry over the oceans. Note, the marked enhancements within the air masses advected from the eastern USA coast toward Europe, and from the eastern Asian coast toward western USA within 20°N and 70°N latitude. Sarwar et al. [2009] already noted a major impact over water bodies. Results exemplify the fact that ozone enhancements are more pronounced over the oceans. Additionally, results over east Asia

![Figure 3. Maximum value of the ratio NOx/VOCs for July August 2004 at ground level.](image)

![Figure 4. Temporal evolution of the daily mean ozone enhancement spatially averaged over Northern hemisphere (NH_mean), Europe (EU_mean), Asia (Asia_mean), North Atlantic Ocean (NA_mean), North Pacific Ocean (NP_mean), and USA (USA_mean) at ground level for July August 2004 due to the inclusion of Crowley and Carl [1997] NO$_2$ photoexcitation chemistry [The location of the averaged areas are plotted in Figure 2].](image)
are in concordance with Li et al. [2011] who reported enhancements over 30 ppbv in major urban areas. Differences between the present results and those from Sarwar et al. [2009], Ensergh et al. [2010], and Li et al. [2011] are mainly attributed to different model resolution (e.g., mesoscale applications compared against global runs), and emission inventories.

[40] Focusing in strong NO$_2$ emission sources, the temporal evolution of the surface ozone concentration in Los Angeles, New York, Paris and Beijing are plotted in Figure 5 for July and August 2004. The hourly concentrations for both the BC and CC are plotted. As noted before, the main differences are observed during peak ozone formation hours. The CC case shows larger ozone concentrations during midday for the two months in all the mega-cities selected. A marked enhancement at noon and a rather slow loss rate of ozone during nighttime are the main impacts of the photoexcited chemistry in strongly polluted areas. Not all days during the two-month time frame show similar enhancements, which is primarily due to meteorological conditions that may lead to strong or weak enhancements during a particular day; i.e., meteorological conditions may sometime dominate gas-phase chemistry, which in turn, can either increase or decrease concentrations of atmospheric species at a given pressure level.

4.1.1 Influence on NO$_2$, HONO and OH Surface Concentrations Compared to O$_3$

[41] The photoexcitation of NO$_2$ promotes the production of HONO and OH radical through pathway described by reaction (12). This new source of NO$_2$ during daytime may contribute to the enhancement of the OH radical and thus the oxidative capacity of the atmosphere, while, also furthering the production of ozone. On the other hand, the NO$_2$* chemistry may contribute to the loss of available NO$_2$ in the atmosphere. The BC and CC model runs are used to compare the enhancements and loss of O$_3$, NO$_2$, HONO, and OH. Figure 6 shows the maximum decreases and enhancement of daily mean O$_3$, NO$_2$, HONO and OH for the period studied.

[42] As stated previously, the CC case exhibits ozone enhancements over the entire globe. However, one can also identify regions where O$_3$ concentrations slightly decrease (Figure 6a). Especially relevant is the result over Australia, where minor O$_3$ enhancements are observed (below 3 ppbv) in addition to non-negligible decreases in O$_3$ concentrations (1–2 ppbv). It is important to note that these decreases are observed during the cold season in the Australian continent. The northern hemisphere also exhibits regions where O$_3$ decreases with the inclusion of Crowley and Carl [1997] kinetics, being the most relevant in northeast India (1–2.5 ppbv). This region is characterized by low NO$_x$ and biogenic emissions, suggesting remote transport of low NO$_x$ air masses.

[43] Figure 6d shows the maximum enhancement of NO$_2$ when comparing the CC and BC runs. Overall increases in NO$_2$ are expected due to the increased oxidation rate associated to reaction (12). Over background regions, the impact of the NO$_2$* chemistry is negligible less than 0.01 ppbv. During the daytime, the photoexcitation chemistry contributes to an overall enhancement of the surface NO$_2$. In areas affected by significant NO$_x$ emissions, the enhancements during the day reach maximum values of 1–6 ppbv. On the other hand, during nighttime, those areas are characterized by a decrease in NO$_2$ concentrations in the CC case. Unlike O$_3$, NO$_2$ variations due to the new kinetics are strongly associated to large NO$_x$ emission sources. In this sense, Figures 6c and 6d show a strong correlation with Figure 1b where NO$_2$ emissions are depicted. The maximum decrease in NO$_2$ is observed in east Asia, where loss of 4 ppbv in NO$_2$ is obtained with the simulations.

[44] The CC case shows an increasing trend in HONO concentrations during most part of the day although the variations are considerably small (Figure 6f). Maximum enhancements and decreases are observed in/near urban locations of 8–9 and 6–7 pptv, respectively. In background regions, the CC simulation shows small HONO enhancements during most of the day, meanwhile the areas influenced by large NO$_x$ emissions present an enhancement during the day and a loss of HONO during the nighttime. The major loss rate is observed during the first daylight hours followed by a sudden enhancement of HONO (not shown). These results indicate that the impact of the NO$_2$* chemistry on NO$_2$ and HONO is negligible.

[45] One of the major results of the photoexcited chemistry is the production of the hydroxyl radical following the direct formation through reaction (12) and through the photolysis of HONO via reaction (4). Figure 6h presents the maximum enhancement of OH for the two months studied. The spatial patterns are rather similar to those associated with O$_3$ impacts (Figure 6b). The areas where the enhancement is more relevant are those located over water bodies and affected by direct NO$_x$ emissions (e.g., ship routes, large industrial areas, urban areas). The maximum enhancement observed in the CC case is around $\sim 6–7 \times 10^{-2}$ pptv ($14–17 \times 10^5$ molec cm$^{-3}$), which are produced during daylight hours. In background continental areas, results show limited

<table>
<thead>
<tr>
<th>Area</th>
<th>Minimum</th>
<th>Average</th>
<th>Maximum</th>
</tr>
</thead>
<tbody>
<tr>
<td>North Pacific</td>
<td>0.8</td>
<td>2.1</td>
<td>2.5</td>
</tr>
<tr>
<td>USA</td>
<td>−3.8</td>
<td>1.9</td>
<td>2.3</td>
</tr>
<tr>
<td>North Atlantic</td>
<td>0.5</td>
<td>2.3</td>
<td>2.6</td>
</tr>
<tr>
<td>Europe</td>
<td>−1.5</td>
<td>1.8</td>
<td>2.0</td>
</tr>
<tr>
<td>Asia</td>
<td>−2.3</td>
<td>1.4</td>
<td>1.6</td>
</tr>
</tbody>
</table>

Table 3. Daily Mean Ozone Enhancement Over Several Areas Around the World Depicted in Figure 2

Table 4. Number of Days With Exceedances of the Enhancement Threshold Values for Hourly Ozone Concentration at Ground Level Over Several Areas Around the World Depicted in Figure 2c

<table>
<thead>
<tr>
<th>Area</th>
<th>Average/Maximum (days)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>≥3 ppbv</td>
</tr>
<tr>
<td>North Pacific</td>
<td>14.5/62</td>
</tr>
<tr>
<td>USA</td>
<td>39.2/62</td>
</tr>
<tr>
<td>North Atlantic</td>
<td>21.6/62</td>
</tr>
<tr>
<td>Europe</td>
<td>32.5/62</td>
</tr>
<tr>
<td>Asia</td>
<td>23.6/62</td>
</tr>
</tbody>
</table>

*Average: number of days exceeding a threshold spatially averaged over an area (over all grids of an area); Maximum: the maximum number of days exceeding a threshold in a specific grid of an area of study.
enhancements during the daytime. Furthermore, it is important to note that in some areas a decrease in the OH radical is also observed (detected during the last daylight hours) (Figure 6g). The ship routes from east USA toward Europe are the regions where the decrease of OH is more remarkable when the photoexcitation chemistry is turned on. The OH radical is extremely reactive, and its concentration in atmospheric conditions range from $8 \times 10^{-13}$ to 0.5 pptv. Thus, the variations observed in the atmospheric equilibrium concentrations are understandably small due to its high reactivity.

4.2. Vertical Distribution of the Ozone Enhancements Due to the NO$_2^*$ Chemistry

Vertical mixing and convective transport are efficient processes to transport surface emission and polluted air masses toward upper layers. The ozone enhancement observed at the ground level is also observed throughout the troposphere. Figure 7 presents the vertical cross sections of the maximum hourly O$_3$ enhancement for July and August 2004 at 90W, 30W, 10E and 130E longitude degrees. As noted in the previous sections, the main impacts are produced in the northern hemisphere, from 10N to 70N latitude degrees. Figures 7a and 7c depicts a well mixed vertical layer within the boundary layer, which contributes to the vertical transport of the O$_3$ enhancement from the surface up to 700 hPa (~3000 m a.s.l.). This vertical mixing is more efficient over continental areas than over water bodies since the vertical distribution over water bodies is due to the advection of well-mixed continental air masses within the PBL and, in some cases, transported beyond it via convective updrafts. The vertical distribution shows an ozone enhancement around 10–20 ppbv in the lower troposphere. Some of the enhancement observed in upper layers is attributed to local surface emissions (i.e., latitudes 50N-60N in Figure 7c, where northern Europe emissions strongly influence the enhanced O$_3$ profile). However, the synoptic transport of enriched air masses from strong emission sources toward the

<table>
<thead>
<tr>
<th>Network</th>
<th>Number of Stations</th>
<th>Observed Mean (ppbv)</th>
<th>Model Mean (ppbv)</th>
<th>MB (ppbv)</th>
<th>RMSE (ppbv)</th>
<th>MNBE (%)</th>
<th>MNGE (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>WDCGG</td>
<td>41</td>
<td>31.7</td>
<td>27.3/30.9</td>
<td>-4.3/-0.8</td>
<td>11.2/10.9</td>
<td>-7.8/4.2</td>
<td>25.5/27.3</td>
</tr>
<tr>
<td>EMEP</td>
<td>70</td>
<td>38.3</td>
<td>27.3/31.1</td>
<td>-11/-7.2</td>
<td>15.8/14.0</td>
<td>-24.2/-13.4</td>
<td>30.4/28.0</td>
</tr>
<tr>
<td>CASTNET</td>
<td>63</td>
<td>31.7</td>
<td>33.8/37.5</td>
<td>2.4/5.8</td>
<td>2.2/10.4</td>
<td>11.6/23.8</td>
<td>22.7/30.1</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Network</th>
<th>Number of Stations</th>
<th>Observed Mean (ppbv)</th>
<th>Model Mean (ppbv)</th>
<th>MB (ppbv)</th>
<th>RMSE (ppbv)</th>
<th>MNBE (%)</th>
<th>MNGE (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>WDCGG</td>
<td>12</td>
<td>1.9</td>
<td>1.2/1.2</td>
<td>-0.7/-0.7</td>
<td>2.0/2.0</td>
<td>29.4/33.2</td>
<td>102.3/102.3</td>
</tr>
<tr>
<td>EMEP</td>
<td>21</td>
<td>3.9</td>
<td>1.8/1.8</td>
<td>-2.1/-2.1</td>
<td>3.4/3.4</td>
<td>38.9/37.9</td>
<td>56.6/55.8</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Network</th>
<th>Number of Stations</th>
<th>Observed Mean (ppbv)</th>
<th>Model Mean (ppbv)</th>
<th>MB (ppbv)</th>
<th>RMSE (ppbv)</th>
<th>MNBE (%)</th>
<th>MNGE (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>WDCGG</td>
<td>14</td>
<td>121.5</td>
<td>145.3/139.8</td>
<td>23.8/18.3</td>
<td>51.5/49.1</td>
<td>43.3/37.8</td>
<td>50.1/47.1</td>
</tr>
</tbody>
</table>

Figure 5. Temporal evolution of the hourly ozone concentration in several mega-cities for July 2004 from BC, CC and LC cases. (a) Los Angeles, (b) New York, (c) Paris, and (d) Beijing.
ocean in the lower and middle troposphere is also a relevant process when assessing the impact of the new chemistry. The cross sections over the Atlantic ocean and the western Pacific clearly depicts a uniform air mass, where the maximum enhancements are 8–9 ppbv. These air masses extend from the surface to about 650 hPa or, in some cases, cover the entire troposphere when approaching the North Pole.

Figure 6. Maximum daily mean (a, c, e, g) decreases and (b, d, f, h) enhancements in near-ground level concentrations for O$_3$ (Figures 6a and 6b), NO$_2$ (Figures 6c and 6d), HONO (Figures 6e and 6f) and OH (Figures 6g and 6h) for July and August 2004 due to the inclusion of Crowley and Carl [1997] NO$_2$ photoexcitation chemistry.
in this intercontinental transport from USA to Europe, and Asia to USA is rather constant within the troposphere. Over the Atlantic (see Figure 7b), a core region, located between 35N and 65N from the surface to 650 hPa shows an enrichment of 8–10 ppbv of ozone. Such air masses strongly contribute to the \( O_3 \) concentrations modeled over Europe.

4.3. Differences in the Impact of Li et al. and Crowley and Carl Kinetic Reaction for the Photoexcitation of NO\(_2\) in Presence of Water Vapor

[45] Given the unresolved differences between past studies that have measured the reaction rate of NO\(_2^+\) with H\(_2\)O and the comparable results of Crowley and Carl [1997] and Carr et al. [2009], the CC and LC model runs are compared for the sake of completeness. Future studies are needed to determine conclusively that NO\(_2\) photoexcitation chemistry indeed occurs under atmospheric conditions. LC and CC cases have been run for July and August 2004. Figure 8a presents the maximum 8-h average ozone difference between cases LC and CC. The first striking result is their similar behavior in background areas. Contrary to initial expectations, the different kinetic rate for reaction (12) of LC and CC, under near-ground level atmospheric conditions, produce similar results in most locations. Major differences, however, appear where large emissions are present. In such places, results show enhancements of 3–6 ppbv in western Europe, western and eastern USA coasts, and Asia. Moreover, in areas with large anthropogenic emissions, like California, northeastern USA, the Baltic sea, the Korean peninsula or China show larger enhancements around 10 ppbv when using the Li et al. [2008] kinetics. Such enhancements are strongly localized in highly polluted areas. Furthermore, Figure 8b shows the 8h-average mean ozone difference between LC and CC for the two months studied. Results show a limited impact when comparing the LC and CC simulations; as stated above, only highly polluted areas show significant mean differences, which remain below 3 ppbv.

[49] Focusing on ozone concentrations in mega-cities, Figure 5 plots the surface ozone temporal evolution of the LC, CC and BC cases. In most situations, LC results behave similar to those produced by the CC case, with an enhancement of several ppbv, compared with the BC. There is no clear difference between results from LC and CC. Namely, the limiting pathway in the photoexcitation of NO\(_2\) is through reaction (10), and the effect of reaction (12) is limited within this context. Under a large variation of atmospheric conditions and emission scenarios, both chemistries, LC and CC, behaves similarly. Only in a few regions, LC produces significant enhancements (below 3 ppbv) when compared to CC. For instance, the South Coast air basin of California presents some enhancements downwind Los Angeles city, while in Los Angeles no relevant differences exist (in agreement with Ensberg et al., 2010 results in Los Angeles, Pomona and Riverside).

Figure 7. Vertical cross sections of the maximum hourly enhancement concentrations for \( O_3 \) at longitude (a) 90W, (b) 30W, (c) 10E, and (d) 130E.
4.4. Evaluation Metrics With and Without the Photoexcited NO\textsubscript{2} Chemistry

[50] All the enhancements discussed in the previous section impact the model performance. Table 5 presents some classical statistics for ground level O\textsubscript{3}, NO\textsubscript{2} and CO from BC and CC model cases. Three different background networks are used. The number of stations per network are provided in the second column of the table. The O\textsubscript{3} model mean for July and August 2004 at surface level increases from 27.3 ppbv to 30.9 ppbv with the CC case when using the WDCGG observation network. Similar trends are obtained when the EMEP and the CASTNET network are used. The largest increase is observed with the CASTNET stations, the model mean increases from 33.9 to 37.5 ppbv. Those model mean increases have a positive impact in the EMEP and WDCGG networks statistics but an increase of the bias and the error is then produced in the CASTNET network.

[51] When NO\textsubscript{2} and CO are evaluated, the effect of the new chemistry produces a rather positive impact on the statistical metrics. Both, bias and error are slightly reduced. The reduction is more significant for CO than NO\textsubscript{2}.

5. Summary and Conclusions

[52] For the first time, the impact of NO\textsubscript{2}* chemistry is assessed at a global scale, during the July August 2004, by
using the new online NMMB/BSC-CTM air quality model. Li et al. [2008] suggested that the reaction of the photoexcited NO₂ with H₂O may be a significant contributing source of OH radicals under atmospheric conditions. They reported a reaction rate for NO₂ with H₂O, which is an order of magnitude greater than what it was previously reported in the literature by Crowley and Carl [1997]. Following the idea of a new OH source from NO₂ with H₂O, several authors have studied the impact on air quality over the USA [e.g., Wennberg and Dabdub, 2008; Sarwar et al., 2009; Ensberg et al., 2010; Li et al., 2011] over USA and China, allowing for a multiscale analysis in other areas also.

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Li, S., J. Matthews, and A. Sinha (2009), Response to comment on “Atmospheric hydroxyl radical production from electronically excited NO2 and H2O2,” *Science*, 324(5925), 336.


